Lab Report Guidelines

Honors Chemistry 2016-17

General Guidelines:

◊ *When* typed you may leave space to insert written calculations. When handwritten you may only write on one side of the paper.

◊ Reports are due two school days after the lab is completed **unless otherwise noted.**

◊ The prelab **must be completed** before the lab or a zero is received, NO EXCEPTIONS.

◊ All sections MUST be clearly labeled (Purpose, Prelab Questions, Procedure…etc) skipping a line between each section.

Prelab 30% *(See lab report RUBRIC for complete description of point distributions):*

Name (presentation pts)

Descriptive Title-(part of presentation points) (Experiment #1 is NOT ACCEPTABLE!)

CLEAR Labels (presentation pts)

Purpose Should consist of: at **minimum** one complete sentence explaining what your objective(s) is/are for the experiment and the method being used to achieve them.

Skip line →

Pre-lab Questions

Answer any assigned pre-lab questions in complete sentences; show all work for any calculations (these may be handwritten). *Skip lines between EACH question!*

Procedure A summary describing how to complete the experiment. *Do not repeat the given procedure step by step.*

Post-Lab Analysis 70%.

Title (presentation pts) Your partners name under your name in parentheses (presentation pts).

Data ALL Data tables must be made using a ruler. **LABEL** all quantities with a unit and report to the proper significant figures. (-0.5 pt for no unit or incorrect significant figures) **All graphs must follow guidelines on graphing guidelines handout for full credit.**

Analysis/Calculations:

◊ Number each step in all calculations (these may be handwritten) and label all numbers with a unit and *identity*, i.e. 0.01 **moles** *NaCl. WATCH SIGNIFICANT FIGURES IN ALL CALCULATIONS!!! Skip lines between each question! Any questions must be answered in complete sentences.*

Conclusion (WHRES) Conclusions restate the purpose and include the results. In 5 complete sentences **MAXIMUM**, address: (W) What did you do (usually purpose)? (H) How did you do it? (Method used, ex: calorimetry…NOT a step by step procedure.) (R) What were your results? (This may include a number(s) but NOT DATA!) **Error analysis:** If applicable, include the % error (E) and 1 possible source (S). Explain how the error source influences the results. Are they higher or lower than they should be?

Technique Follow proper procedures and directions. Points are deducted at the instructor's discretion.

Example Conclusion:

The mass of water was determined (what?) using the density formula, the measured volume of a sample of water and the literature density value at the given temperature (how?). The results according to the density formula are that the mass of the water sample is 10.85g (results, NOT DATA). This represents a 1.36% error (error) from the measured value of 11.00g. The experimental mass may have been higher due to water droplets inside the cylinder that added to the mass but were not measured in the volume (source of error). *Some lab reports may exclude some of the sections. You will be informed when this occurs.